46,002 g I2 = 46,002/253,8 mol I2 = 0,1813 mol I2

1,011 g H2 = 1,011/2,016 mol H2 = 0,5015 mol H2

Over 1,902 g I2 = 1,902/253,8 = 7,494.10-3 mol I2

0,1813-7,494.10-3 mol I2 is er omgezet = 0,1738 mol I2

Nu tabel gebruiken

|  |  |  |  |
| --- | --- | --- | --- |
|  | H2 | I2 | HI |
| begin | 0,5015 mol | 0,1813 mol |  - |
| Omzetting | * 0,1738
 | * 0,1738 mol
 | +0,3476 |
| Eind | 0,3277 | 7,494.10-3 | 0,3476 |

 [HI]2

K = -------------

 [H2].[I2]

Nu heb je dus het volume nodig om de concentraties te berekenen.

 (pHI)2

Kp = -----------

 pH2 . pI2